## **INDIAN SCHOOL MUSCAT**

**SET A** 

## **FINAL TERM EXAMINATION**

## **NOVEMBER 2018**

## **CLASS X**

Marking Scheme – SCIENCE [THEORY]

	VALVUE POINTS	marks				
	SECTION A					
1.	BIO	1				
2.	Receptors (1/2) Sense organs (1/2)	1				
	SECTION B					
3.	a) Calcium sulphate hemi hydrate 1M CaSO <sub>4</sub> .1/2H <sub>2</sub> O	2				
4.	b) CaSO <sub>4</sub> .1/2 H <sub>2</sub> O + 3/2 H <sub>2</sub> O → CaSO <sub>4</sub> . 2H <sub>2</sub> O 1M Cover with glass plate (1 mark) Replace plane mirror with concave mirror(1 mark) OR	2				
5.	Two advantage of hydro energy (1mark) Two disadvantage (1mark) Evolution (1), ear lobe, skin clour, or any (1)	2				
	SECTION C					
6.	Method to harness ocean thermal energy.	3				
	One limitation (1/2)					
	One advantage (1 /2 )					
7.	Fuse is used to protect appliances due to short circuit or overloading. (2)	3				
	When excess current flows through the circuit fuse wire gets heated up and melt.					
	If a fuse is replaced by one with larger rating , the device may get damaged while protecting fuse doesn't burn off. (1)					
8.	$h_o = 4.5 \text{ cm}$	3				

u = -12 cmf = +15 cm1/u + 1/v = 1/f $(\frac{1}{2} \text{ m})$ 1/v = 1/f - 1/uv = 60/9 = 6.7 cm (1) m = -v/u $(\frac{1}{2} \text{ m})$ =-6.7/-12=0.558 (1m) OR  $H_0 = 5\ cm$ F = -10 cmu = -20 cm1/f = 1/v - 1/u $(\frac{1}{2})$ V = -6.67 cm  $(\frac{1}{2})$ Virtual, erect Diminished (1)  $h_i/h_o = v/u$  $(\frac{1}{2})$  $h_{i} = 1.67$  cm  $(\frac{1}{2})$ Ray diagram object between F and P of concave mirror. (2m)3 Virtual, Enlarged (1m)a)  $1^{\text{st}}$  element has 1 valence electron and last element has 8 V.E. 3 b) Hydrogen resembles with both; group 1 (alkali metals) and group 17 (Halogens). Therefore occupies a unique position. Or any other relevant answer. c) Very inert, extremely low concentration... (3x1=3)a) Electricity is carried through ions.. Hydrogen ions in acids are produced only in presence of 3 water. b)HCl ionizes completely in water and produces large amount of H+ ions whereas NH<sub>4</sub>OH ionizes partially in water producing less amount of OH ions. c) Chlorine at anode...Hydrogen at cathode....NaOH near the cathode. (3x1=3)a) Acid is strong enough to attack the enamel of our teeth and corrode it. b) Correct definition...

9.

10.

11.

	On heating, blue copper sulphate crystals turn white. ½ M Tiny droplets of water are seen in the boiling tube½ M	
12.	a)Electron dot structure $\frac{1}{2}$ M, Ion formation1/2 M b) Oxides which react with acids as well as bases to form salts and water. Eg. Al <sub>2</sub> O <sub>3</sub> , ZnO etc. $\frac{1}{2}$ x 2 = 1M	3
13.	c) Due to strong forces of attraction between ions. 1M  Xylem vessels and tracheids form a continuous channels(1/2)  Roots takes up ions actively (1/2)  Ionic difference between soil and water is compensated by absorption of water by roots (1)  Transpiration pull or suction force pull water from the xylem vessels (1)	3
	OR Transport of soluble product of photosynthesis through phloem(1/2) It take place through sieve tube with the help of companion cells (1/2) It take place upward and down ward direction (1/2) Sucrose is transported into the phloem tissue utilizing energy from ATP (1/2) As the osmotic pressure of phloem tissue increases which is compensated by the intake of more water (1/2) This osmotic pressure permits movement of materials in the phloem from high pressure to low pressure (1/2)	
14.	a) Green (1) b) 75% (1)	3
	c) 3:1 (1)	
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	SECTION D	
16.	Myopia (1)	5
	Ray diagram showing defect of myopia (2)	
	Causes of myopia ( $\frac{1}{2} + \frac{1}{2}$ )	
	Concave lens (1)	

OR

	Dispersion -defenition (1)				
	Violet, Red (1)				
	Cause of dispersion (1)				
	Ray diagram for dispersion(2)				
17.	Voltmeter (1m)	5			
	In parallel				
Since high resistance device(1m)					
Defenition of Volt (1 m)  Symbol of variable resisitor (1)					
18.	i) $2,8,2$ ii) $V.E=2$ , Shells = 3 iii) $H>G>F>E$ iv) Reactivity increases	5			
19.	v) A <sub>2</sub> O (1x5=5) (A) Cinnabar (1M)	5			
20.	(i) $2\text{HgS} + 3\text{O}_2$ Heat $\rightarrow 2\text{HgO} + 2\text{SO}_2$ (ii) $2\text{HgO}$ Heat $\rightarrow 2\text{Hg} + \text{O}_2$ (2M) (B) Electrolytic Reduction (1M) (C) Lead and TinLow Melting Point ( $\frac{1}{2}$ X2=1M)				
21.	<ol> <li>a) Pepsin – gastric gland in the stomach, protein digestion         Amylase – salivary gland in the mouth or pancreas in the stomach, digestion of starch into maltose or sugar (2)</li> <li>b) Largest part, villi, plenty of capillary system or any two(2)</li> <li>c) Protect the inner wall of the stomach from the action of acid (1)         OR</li> <li>a) Diagram (1) four labeling (2)</li> <li>b) Function of the four parts (2)</li> </ol>				
	SECTION E				

22. PHY 2

23.	PHY		2
24.	A) Highly acidic – Red /PinkHighly alkaline – Deep blue /Violet B) Weekly acidic – Yellowish greenWeekly basic – Greenish blue OR	1M 1M	2
	i) A = Acidic ,B = Basic 1M ii) greenish blu	ie 1M	
25.	Double displacement 1M BaSO <sub>4</sub> 1M		2
26.	Budding (1)		2
	Diagram in correct sequence (1)		
27.	a) Transpiration or respiration or any two (1 mark each)		2
	OR		
	a) No respiration and no evolution of CO <sub>2</sub> (1)		
	b) Absorption of CO <sub>2</sub> (1)		